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1. (w = 25 m) $(H_2^1 O^8)$

(2. Mr_1: 60г; 300·5 = 1500; Mr_2 = 40 гр)

4. Шешуі: $Mr(HNO_4)$ = 1 + 14 + 16·4 = 79 гр
$Mr(AgNO_4)$ = 108 + 78 = 185 гр

79 + 185 = 264 - $3{,}84 \cdot 10^{24}$

$W = \frac{100 \cdot 40{,}506\%}{30{,}38\%} = 133{,}33\%$

(2. Mr_x ?)

3. Mr(MgCl) = 24+35 = 59

ZnCl = 70+35 = 100

H_2O = 1·2+16 = 18

$CuCl_2$ = 63+35·2 = 133

AgNa : (10.8) 108 + 23 = 131

1. $\omega(NH_3)$ = 20%

$m(NH_3)$ ерім) = 80%

$W = \frac{m(ер.з)}{m(ер.и)} \cdot 100\%$

$20\% = \frac{x}{80} = 100\%$

x = 16 NH_3

$m(H_2O)$ = 80 - 16 = 64

$NH_3 + H_2O \rightarrow NH_4OH$

NH_4OH — 1 N атом N - 7ē
— 5 H атом H - 5ē — 20 ē
1 O атом O - 8ē

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2. $M_r(C_2H_4O_2) = 12 \cdot 2 + 1 \cdot 4 + 16 \cdot 2 = 60$ гр

$\omega = 10\%$

2. $M_{rt} = 60$ гр $300 \cdot 5 = 1500$ $60 - 3 = 57$.

$M_r = 40$ гр