

1. ($w = 25\%$) ($H_2^{\frac{1}{2}}O^{\frac{8}{8}}$)

2. $Mr_1 = 60$; $300 \cdot 5 = 1500$; $Mr_2 = 40$ зр)

4. Мәшһүгі: $Mr(HNO_3) = 1 + 14 + 16 \cdot 3 = 61$ зр
 $Mr(AgNO_3) = 108 + 14 + 16 \cdot 3 = 170$ зр

$79 + 185 = 264 = 3,84 \cdot 10^{24}$

$w = \frac{100 \cdot 40,506\%}{30,389\%} = 133,33\%$

(2. Mr_x)

3. $Mr(MgCl) = 24 + 35 = 59$

$ZnCl = 70 + 35 = 105$

$H_2O = 1 \cdot 2 + 16 = 18$

$CuCl_2 = 63 + 35 \cdot 2 = 133$

$AgNO_3 = (10 \cdot 8) 108 + 23 = 132$

1. $w(NH_3) = 20\%$

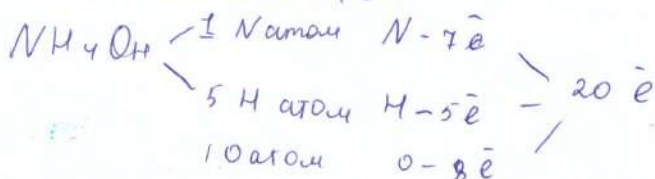
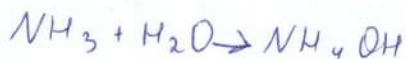
$m(NH_3) \text{ ерім} = 80\%$

$w = \frac{m(\text{ер.3})}{m(\text{ер.4})} \cdot 100\%$

$20\% = \frac{x}{80} = 100\%$

$x = 16 NH_3$

$m(H_2O) = 80 - 16 = 64$



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$$2. M_r(C_2H_4O_2) = 12 \cdot 2 + 1 \cdot 4 + 16 \cdot 2 = 60 \text{ ұр}$$

$$w = 10\%$$

$$2. M_r t = 60 \text{ ұр} \quad 300 \cdot 5 = 1500 \quad 60 - 3 = 57.$$

$$M_r = 40 \text{ ұр}$$